



UCLA

# Stereoselective Green Chemistry Strategies Using Crystal-to-Crystal Reactions: The Advantages of Molecular Nanocrystals

*(Syntheses of complex structures with adjacent quaternary stereogenic centers)*



C (N) S I

*National Science  
Foundation*

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**MCTP**  
Materials Creation Training Program



# Nanotechnology

## —*Potential Benefits:*

Drug delivery, solar energy conversion, catalysis, microelectronics, cosmetics, water and soil remediation, etc.

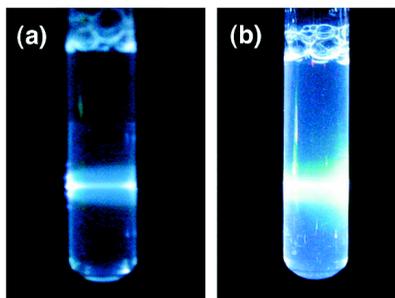
## —*Concerns:*

Toxicology, bioaccumulation, oxidative stress (ROS), transport of pollutants, etc.

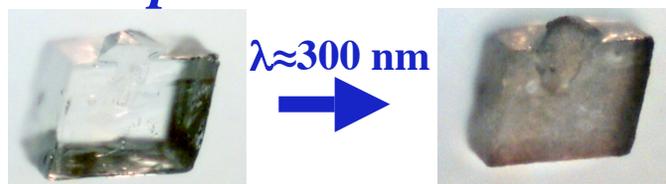
## —*Unknown:*

????? Incentives for New Chemical Reactivity Paradigms (challenge your perspectives on chemical synthesis)

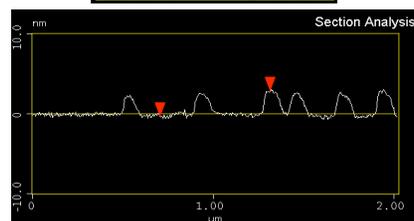
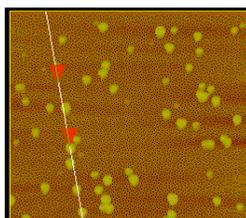
# Challenging or Complex Chemical Processes in Molecular Crystals



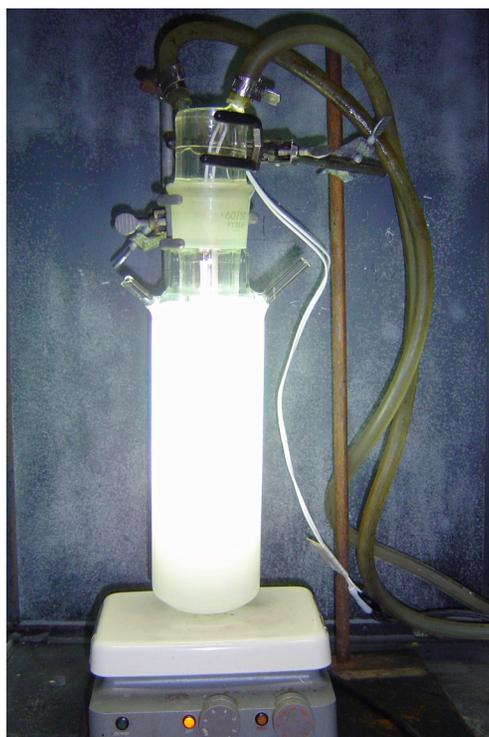
Loading: (a) 0.015 and (b) 1.8 mg/ml



100 % selectivity, but slow and inefficient (limited conversion)



Loading: 20 g / 5 L



Loading: 2 g / 200 ml

Further scale up



Collect product by filtration, recycle water.

# *Pharmaceuticals (complex structures)*

## **The top drugs worldwide (3000-4000 Tons/Year)\***

### **Top 100**

- **\$85.2 billion**
- **50 single enantiomers, sales of \$42.8 billion**

### **Top 300**

- **\$124.4 billion**
- **158 single enantiomers, sales of \$64.7 billion**

### **Top 500 drugs**

- **\$135.9 billion total**
- **269 single enantiomers, sales of \$71.1 billion.**

*Enantiomers: Chemical structures that differ only by the 3D atomic arrangement and have a relation like that of our hands.*

*(Enzymes, DNA, sugars, biological receptors etc. are single enantiomers)*

*\*Chem. Eng. News 1997*

# Lundbeck Pharmaceuticals (1999)

*(chemical complexity has a tremendous cost!)*

## Raw materials

Organic solvents: 990 tonnes  
Other raw materials: 1,640 tonnes

## Water

Crude water: 14,800 m<sup>3</sup>  
Potable water: 7,900 m<sup>3</sup>

## Energy

Oil: 4,450 MWh  
Electricity: 3,300 MWh



## Wastewater:

Wastewater: 14,200 m<sup>3</sup>  
Organic solvents: 0.3 tonnes  
Organic matter: 36 tonnes

## Atmospheric emissions

Organic solvents: 18 tonnes

## Waste

Chemical waste: 1,255 tonnes

Produced tonnes: 179

$$\begin{aligned} \text{E factor} &= \text{Materials Used} / \text{Product} \\ &= 14.7 \end{aligned}$$

# Major Sources of Waste

## —Stoichiometric Acids and Bases

( $\text{H}_2\text{SO}_4$ ,  $\text{AlCl}_3$ ,  $\text{ZnCl}_2$ ,  $\text{BF}_3$ ,  $\text{NaOH}$ ,  $\text{NaOMe}$ , etc.)

## —Stoichiometric Oxidants and Reductants

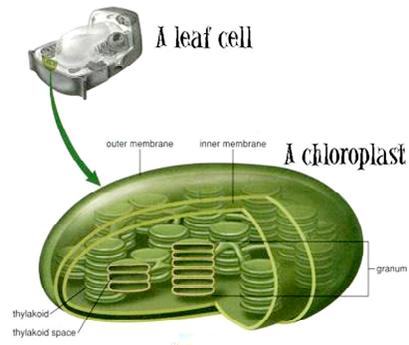
( $\text{Na}_2\text{Cr}_2\text{O}_7$ ,  $\text{KMnO}_4$ ,  $\text{MnO}_2$ ,  $\text{LiAlH}_4$ ,  $\text{NaBH}_4$ ,  $\text{Zn}$ ,  $\text{Fe/HCl}$ , etc)

## —Solvent Losses

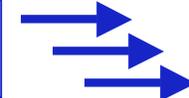
(air emissions and aqueous effluents)

	Ton	E-Factor
—Bulk Chemicals	$10^4$ - $10^6$	< 1 - 5
—Fine Chemicals	$10^2$ - $10^4$	5 - > 50
—Pharmaceuticals*	10 - $10^3$	25 - > 10

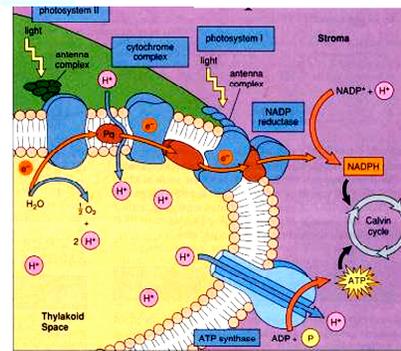
*\*Many are Derived from Complex Natural Products*



# Photosynthesis

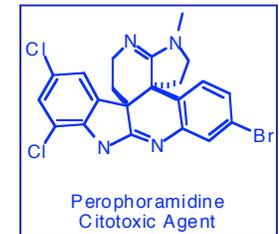
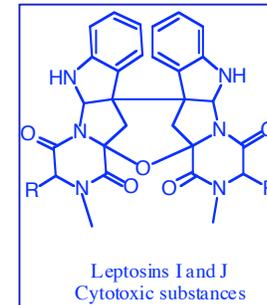
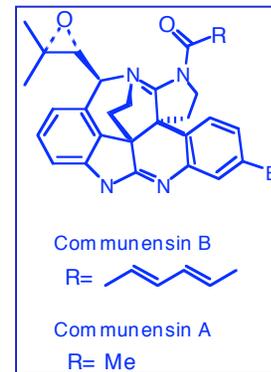
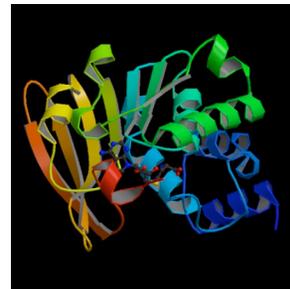
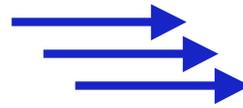


Life



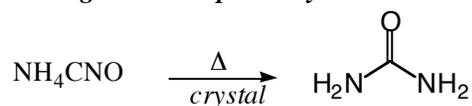
# Biosynthesis

*isoprenes*  
*propionates*  
*sugars*  
*aminoacids,*  
*etc...*

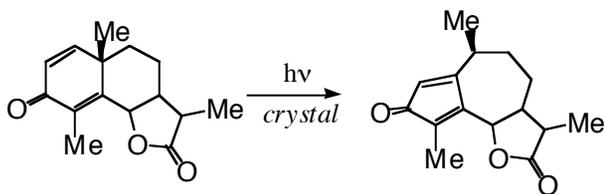


*Secondary metabolites*

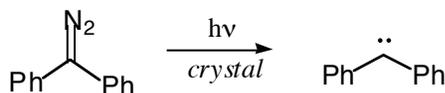
Wholer 1828: First "organic" compound synthesized



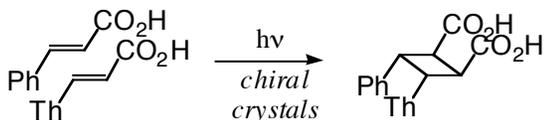
Pfizer 1849: First drug formulated in the US



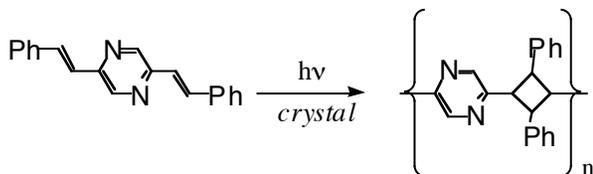
Murray et al. 1962: First triplet molecule observed



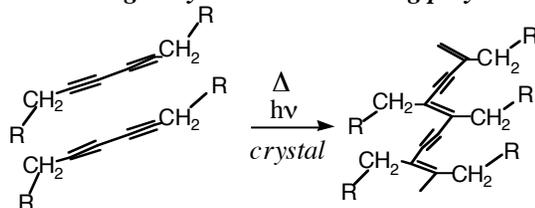
Schmidt et al. 1971: "Spontaneous" generation of optical activity



Hasegawa et al. 1972 : Single crystalline polymers



Wegner et al. 1971: Single crystalline conducting polymers



## REACTIONS IN CRYSTALS

- Associations (components)
- Rearrangements
- Fragmentations

*Have selectivities, specificities and efficiencies that rival those observed in enzymatic process*

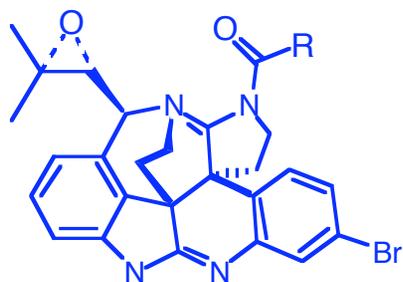
**\*\*\*DPCP\*\*\***

*How to take advantage of  
Photo-[Complex Organic]-Synthesis?*

**Where to start...? Need a Niche!**

# Complex Natural Product Synthesis

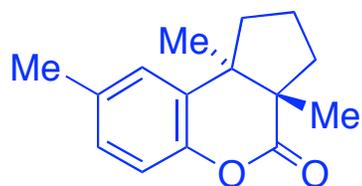
—Compounds with adjacent stereogenic quaternary stereogenic centers



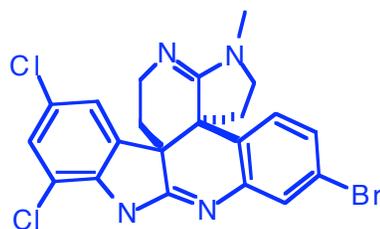
Communensin B



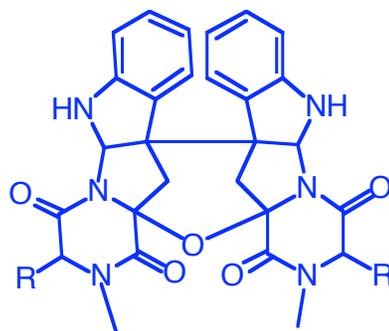
Communensin A



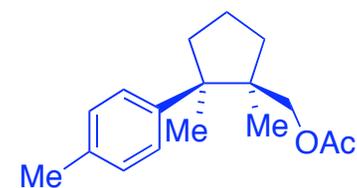
Herbertenolide  
Cytotoxic Agent



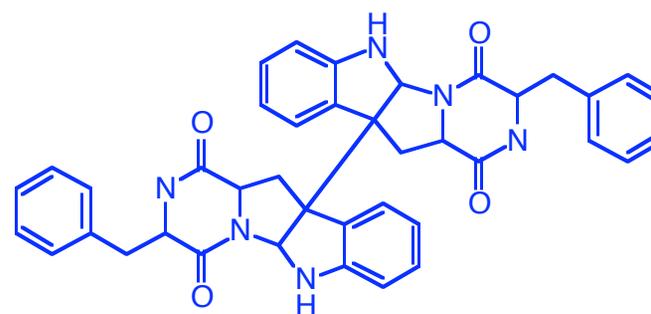
Perophoramidine  
Cytotoxic Agent



Leptosins I and J  
Cytotoxic substances



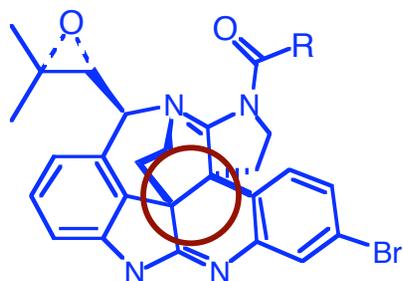
Tochuinyl Acetate  
Cytotoxic Agent



Psycholein  
Antagonists to cholecystokinin  
(CCK)B/gastrin receptor

# Complex Natural Product Synthesis

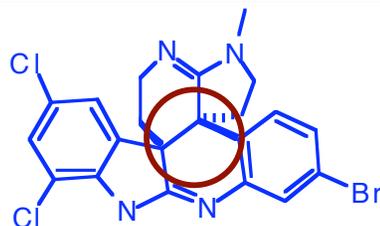
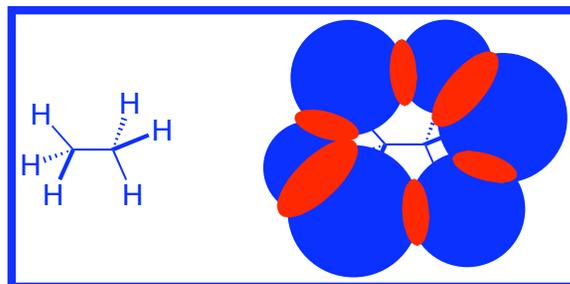
—Compounds with adjacent stereogenic quaternary stereogenic centers



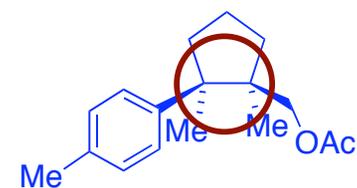
Communensin B



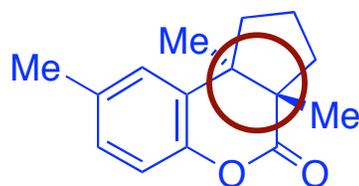
Communensin A



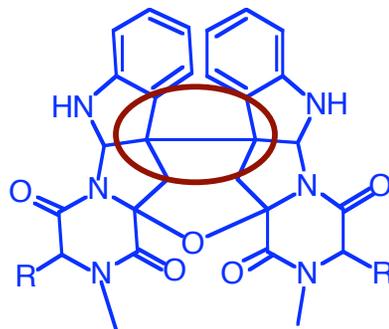
Perophoramidine  
Cytotoxic Agent



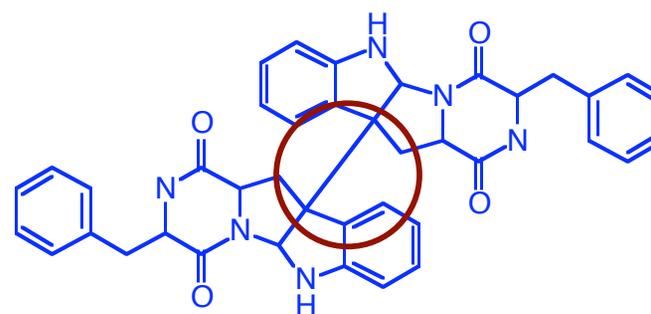
Tochuinyl Acetate  
Cytotoxic Agent



Herbertenolide  
Cytotoxic Agent



Leptosins I and J  
Cytotoxic substances

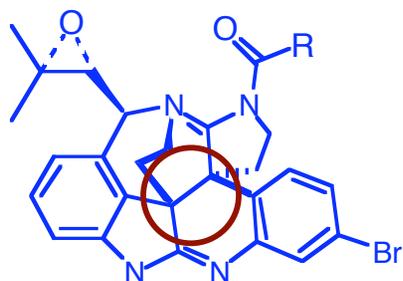


Psycholein  
Antagonists to cholecystokinin  
(CCK)B/gastrin receptor

- 1) *Multiple steps*
- 2) *Poor stereochemical control*
- 3) *Difficult purification*
- 4) *Low yields*

# Complex Natural Product Synthesis

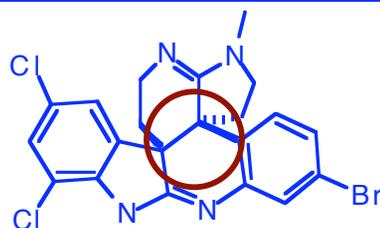
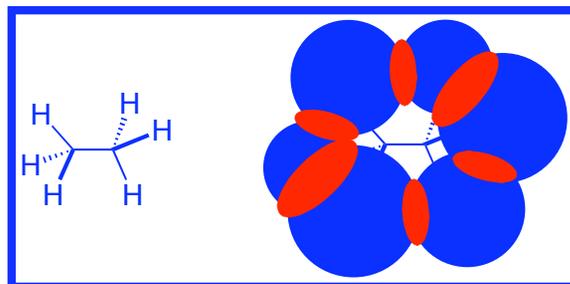
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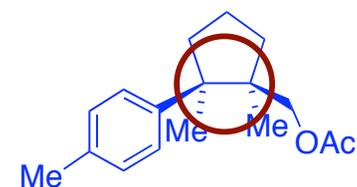
Communensin B



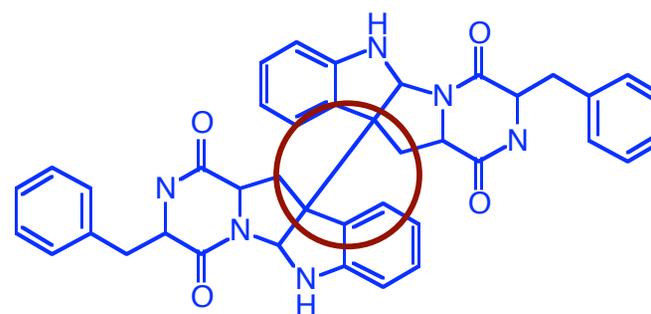
Communensin A



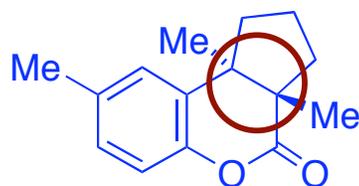
Perophoramidine  
Cytotoxic Agent



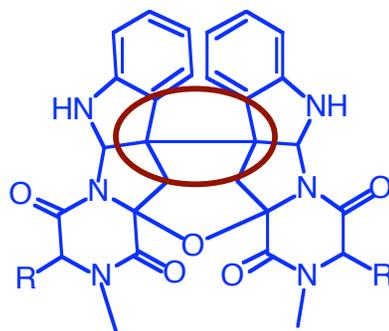
Tochuinyl Acetate  
Cytotoxic Agent



Psycholein  
Antagonists to cholecystokinin  
(CCK)B/gastrin receptor



Herbertenolide  
Cytotoxic Agent

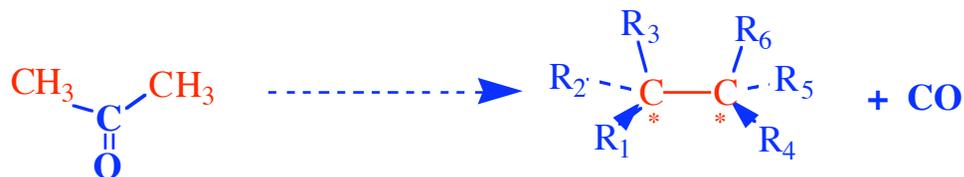


Leptosins I and J  
Cytotoxic substances

E. J. Corey

“The asymmetric construction of molecules with adjacent quaternary centers represents a very challenging area in organic synthesis” *Angew. Chem. Int. Ed.* 1998, 37, 388-401

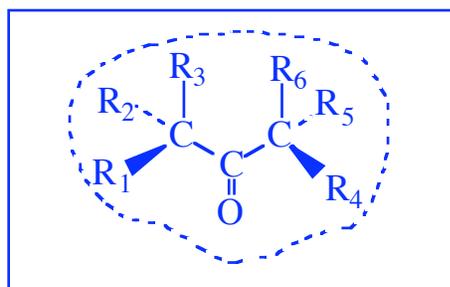
# Solvent-Free Syntheses of Adjacent Stereogenic Quaternary Centers: General Strategy



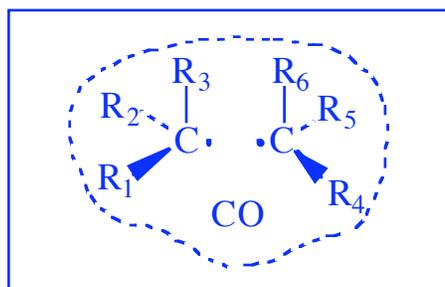
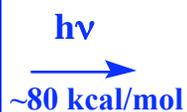
a) Stereoselective ketone chemistry

b) Crystallization

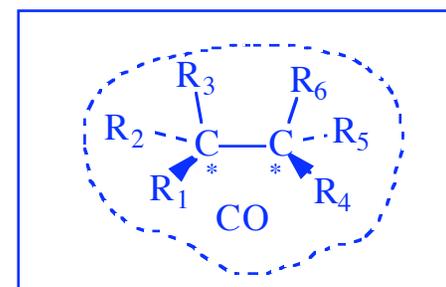
- No solvent
- No separation
- No purification
- Works w/Sunlight



*Reactant Crystal*



*Close neighbors boundaries: ( - - - - )*



*Product Crystal*

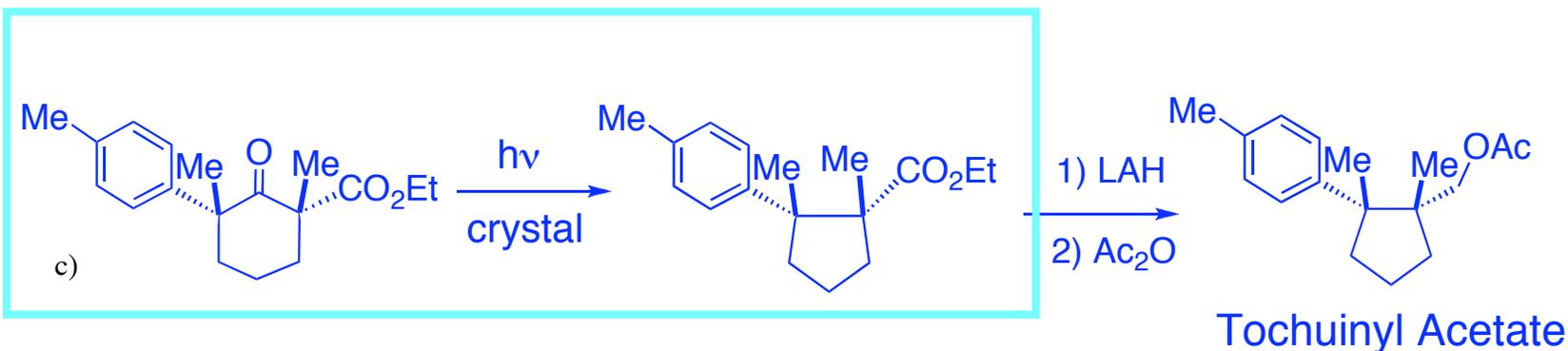
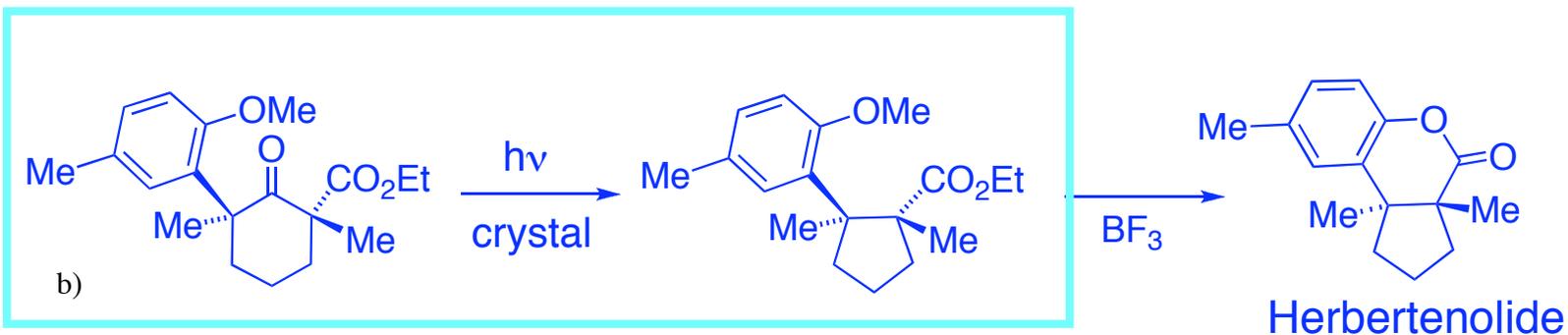
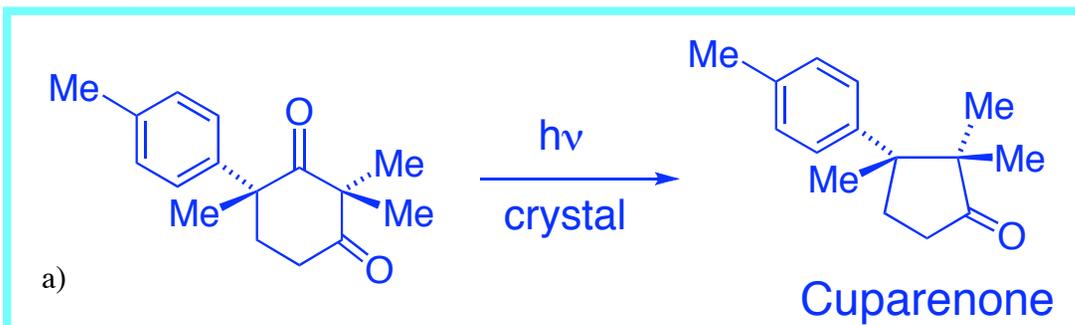


- Highly Predictable Reaction
- Rigidity and Homogeneity lead to:

a) Stereoselectivity

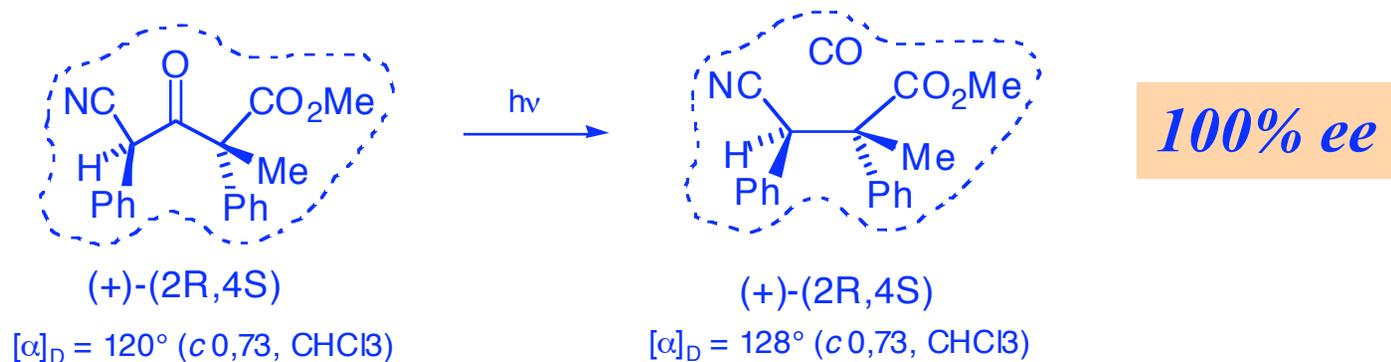
b) Stereospecificity

# Natural Product Syntheses Completed with Solid-to-Solid Reactions

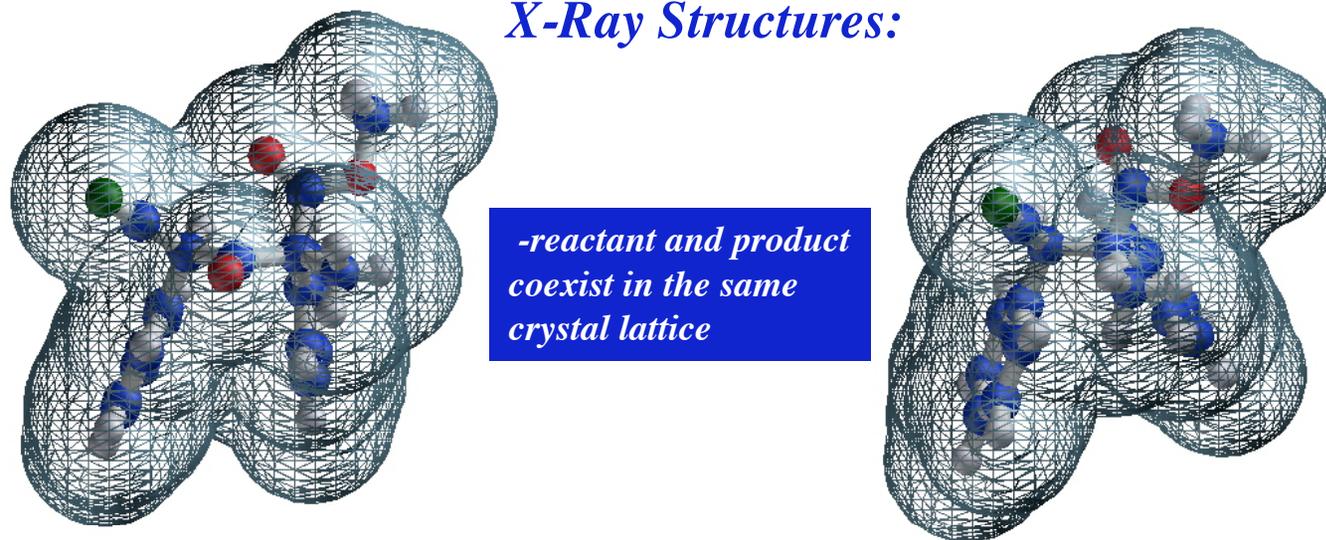


a) Arankumar et al., *Angew. Chem.* **2007**, b) Ng, D.; Yang, Z.; Garcia-Garibay, M. A., *Org. Letters*, **2004**, 6, 645, d) Yang et al., in preparation

*Reaction Stereoselectivity is determined by the similarity between the structures of the reactant and product*

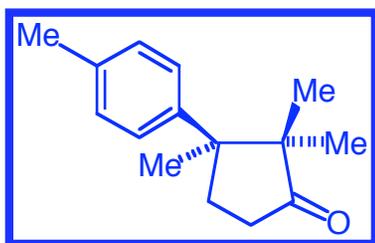


*X-Ray Structures:*

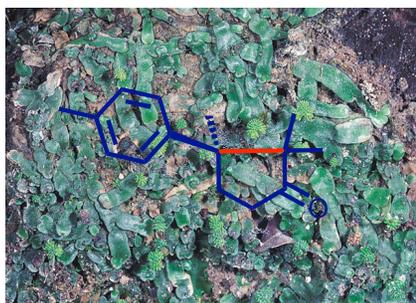


*Crystal Rigidity*  $\Rightarrow$  *Reaction selectivity*  
*Crystal Homogeneity*  $\Rightarrow$  *Reaction Specificity*

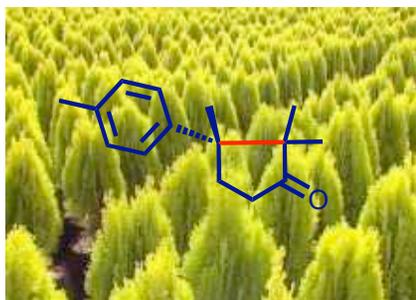
## Reported Syntheses of Cuparenone



5 steps: 2 “pots” and  
one solid state reaction  
60% total yield, both  
enantiomers obtained  
(key step ~100 yield)



(S)-(+)- $\alpha$ -Cuparenone  
“Mannia Fragrans” Liverwort

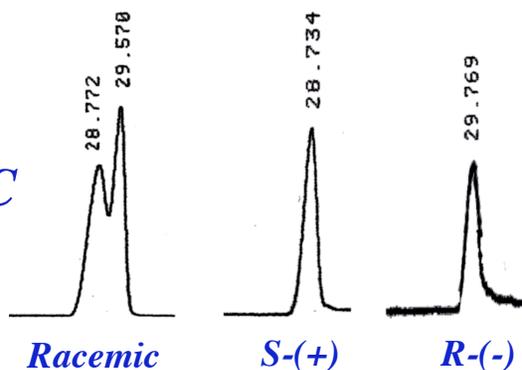


“Mayur pankhi” Tree  
(R)-(-)- $\alpha$ -Cuparenone

Literature	No. of steps	Overall Yield (%) [%ee] <sup>§</sup>
Posner <i>et al.</i> , <sup>16</sup> 1984	8	5 [71]
Taber <i>et al.</i> , <sup>17</sup> 1985	7	0.5 [96]
Kametani <i>et al.</i> , <sup>18</sup> 1985	14	5 [9]
Meyers <i>et al.</i> , <sup>19</sup> 1986	8	29 [97.6]
Greene <i>et al.</i> , <sup>20</sup> 1987	7	18 [99]
Asaoka <i>et al.</i> , <sup>21</sup> 1988	12	1.5
Takano, <i>et al.</i> , <sup>22</sup> 1989	13	0.5
Gharpure <i>et al.</i> , <sup>23</sup> 1989	6.5	6.5 [96]
Fadel <i>et al.</i> , <sup>24</sup> 1991	11	22
Nemoto <i>et al.</i> , <sup>25</sup> 1992	9	3
Canet <i>et al.</i> , <sup>26</sup> 1992	12	12
Honda <i>et al.</i> , <sup>27</sup> 1993	9	6
Maruoka <i>et al.</i> , <sup>28</sup> 1996	7	10 [70]
Kosaka <i>et al.</i> , <sup>29</sup> 1997	14	3.6
Nakashima <i>et al.</i> , <sup>30</sup> 2000	17	5.6
Satoh <i>et al.</i> , <sup>31</sup> 2003	9	32
Spino <i>et al.</i> , <sup>32</sup> 2004	9	12
Our work	5	60 [99]

<sup>§</sup> Only %ee values ex

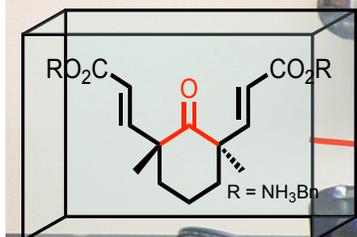
Chiral GC



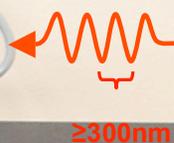
# *Why is it taking so long to develop solid state chemistry ?*

- There was no incentives !*
- It is possible to “engineer” reactions in crystals*
- Most problems in scale up and optimization seem surmountable ... with nanocrystals*

# Reaction Monitoring by Cross-Polarizing Microscopy

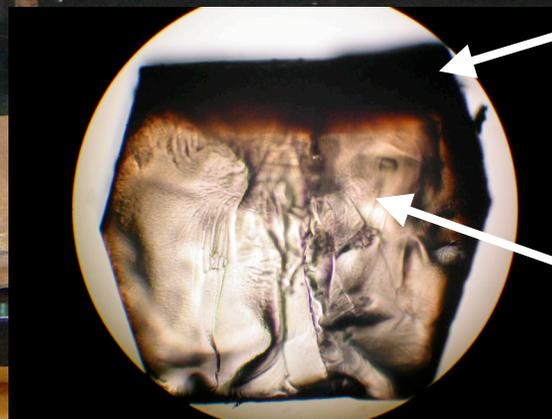


UV  
Light

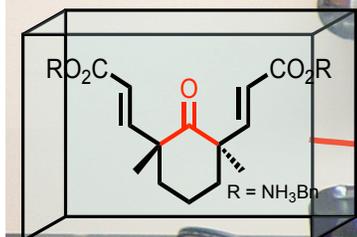


100% rxn

No rxn



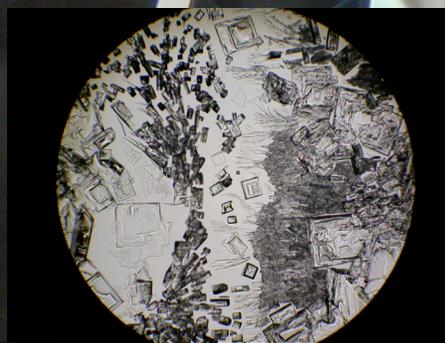
# Reaction Monitoring by Cross-Polarizing Microscopy



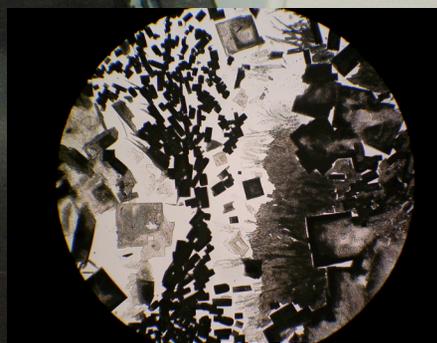
UV  
Light



≥300nm

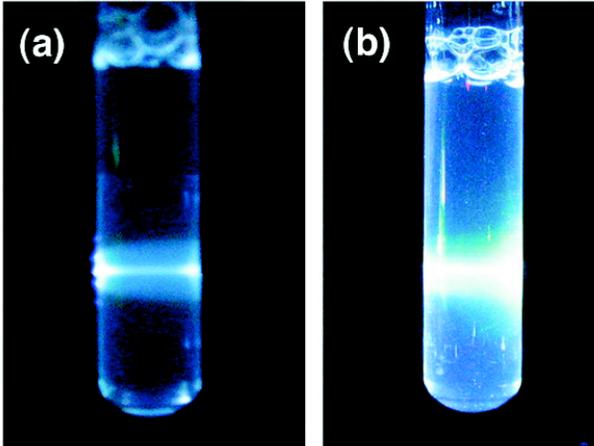


97% rxn

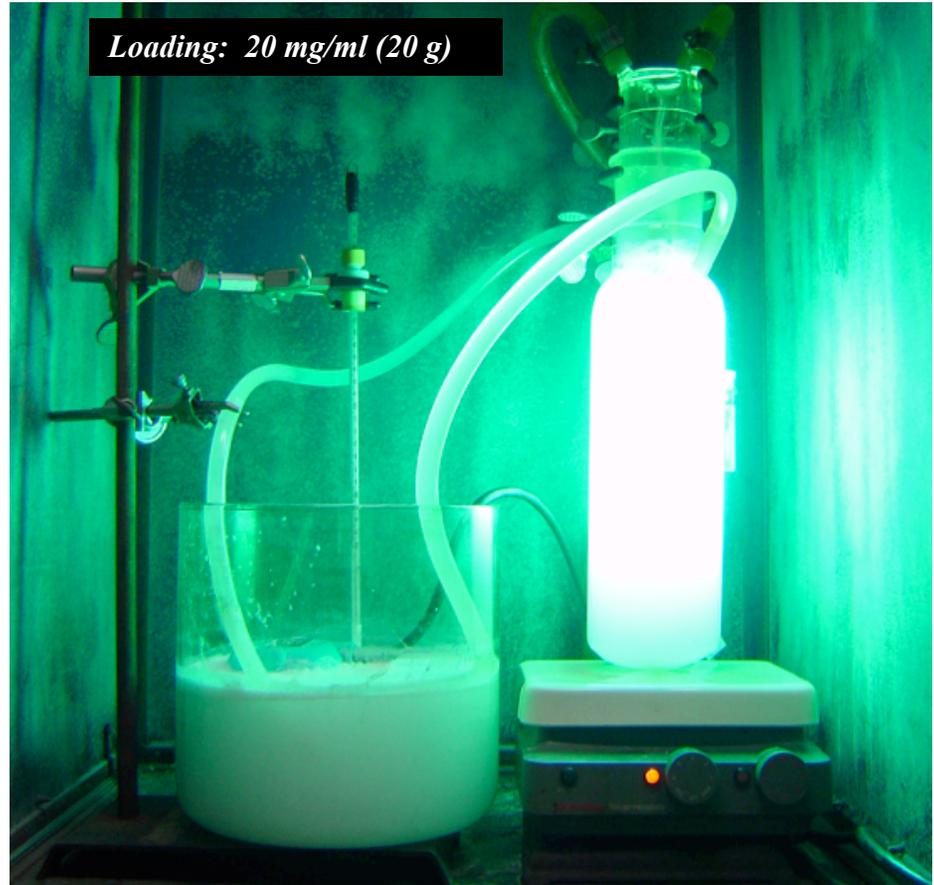


# *Reduce the Crystal Size!*

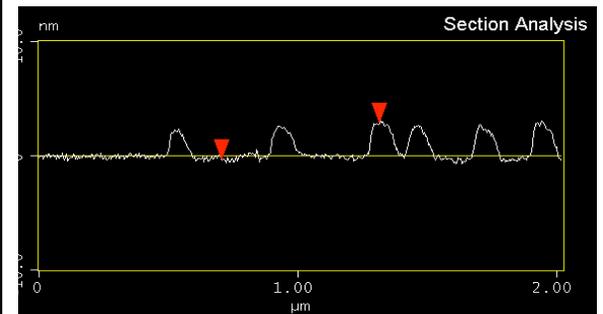
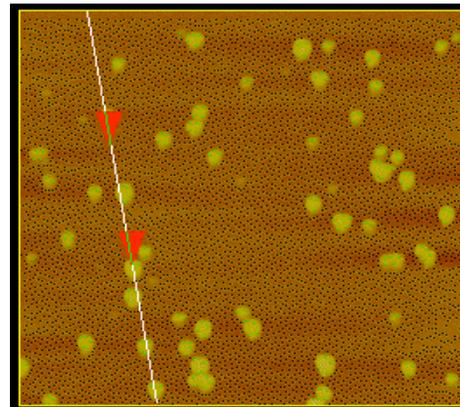
- 1. Nanocrystals of 100-500 nm retain the chemical influences of bulk crystals*
- 2. As particle sizes approach the wavelength of individual photons : They are excited in nearly homogeneous manner*
- 3. With strongly reduced birefringence, the interaction between particles and photons will approach those of molecules and photons*
- 4. Small particles form stable suspensions that can be transported by fluids (H<sub>2</sub>O)*
- 5. With reduced birefringence, one can use transmission spectroscopy and monitor in situ*
- 6. Absolute quantum and chemical yield measurements possible*
- 7. Free-flowing suspensions allow the easy scale up of solid state photochemical reactions by taking advantage of flow reactors*
- 8. With suspensions made in water, the use of solid state chemistry will allow for the synthesis of complex structures by very desirable Green Chemistry methods*



*Loading: (a) 0.015 and (b) 1.8 mg/ml*



*Loading: 20 mg/ml (4 g)*





## Reactions in Crystals:

- Highly Selective

- No Solvents

- No Purification

- .

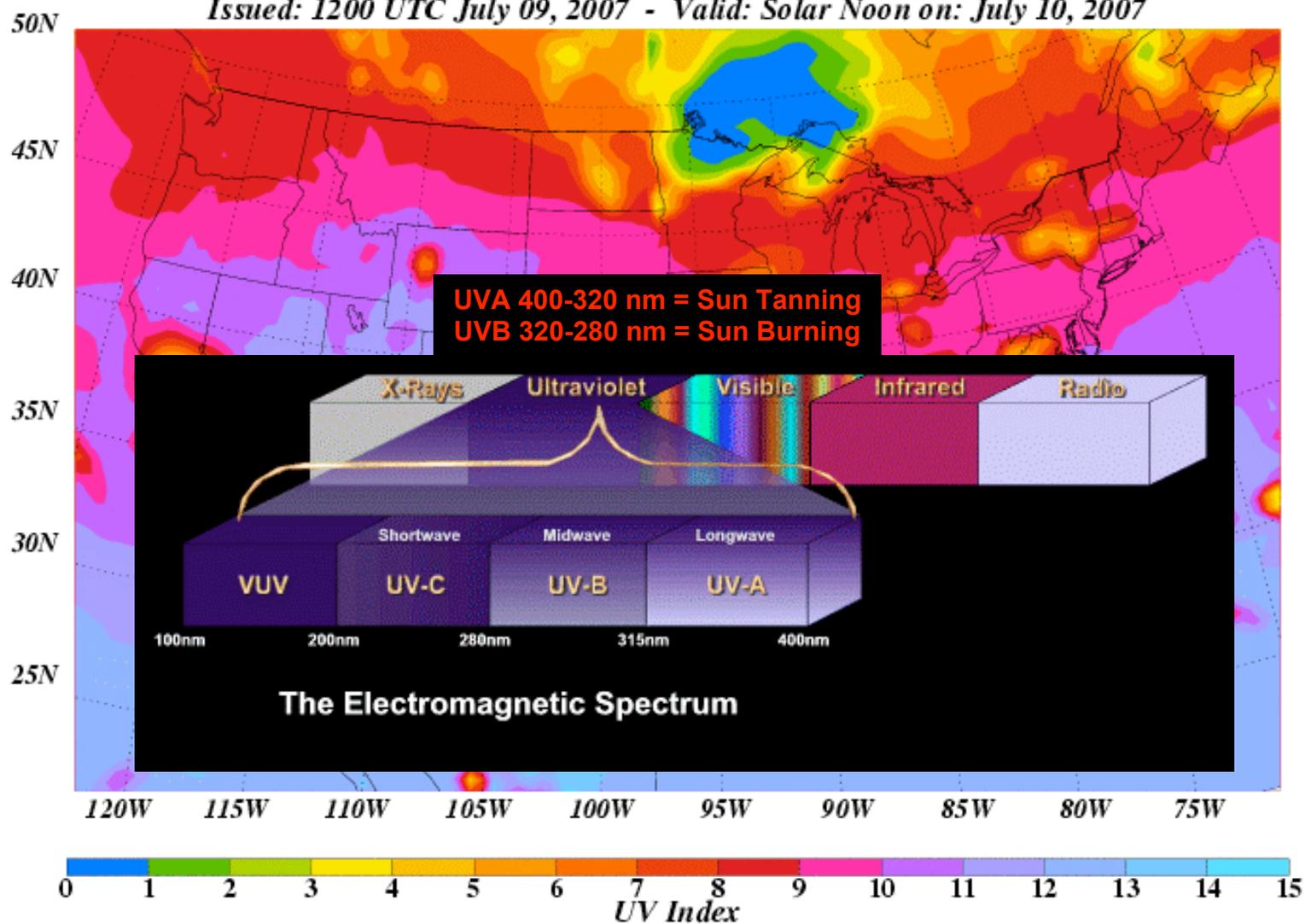
- .

- .

Can one use sunlight?

# UV INDEX FORECAST

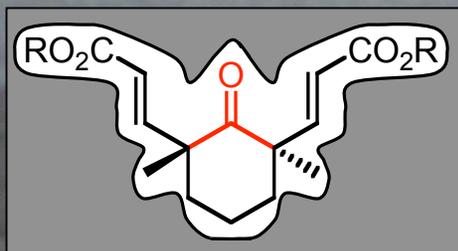
Issued: 1200 UTC July 09, 2007 - Valid: Solar Noon on: July 10, 2007



*How efficient is decarbonylation in sunlight?*

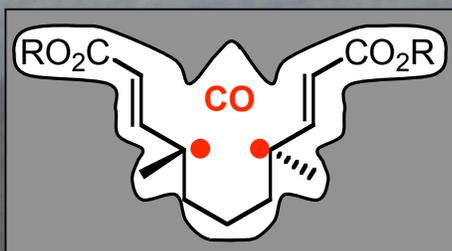


*~2h exposure by noon in January in LA*



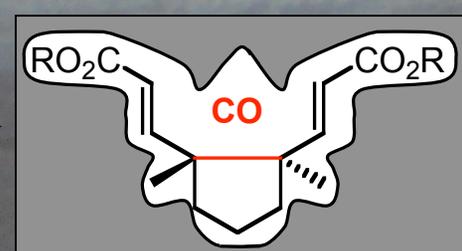
R = NH<sub>3</sub>Bn

- 1)  $h\nu$
- 2) ISC
- 3)  $\alpha$ -cleav
- 4) -CO



CRYSTAL

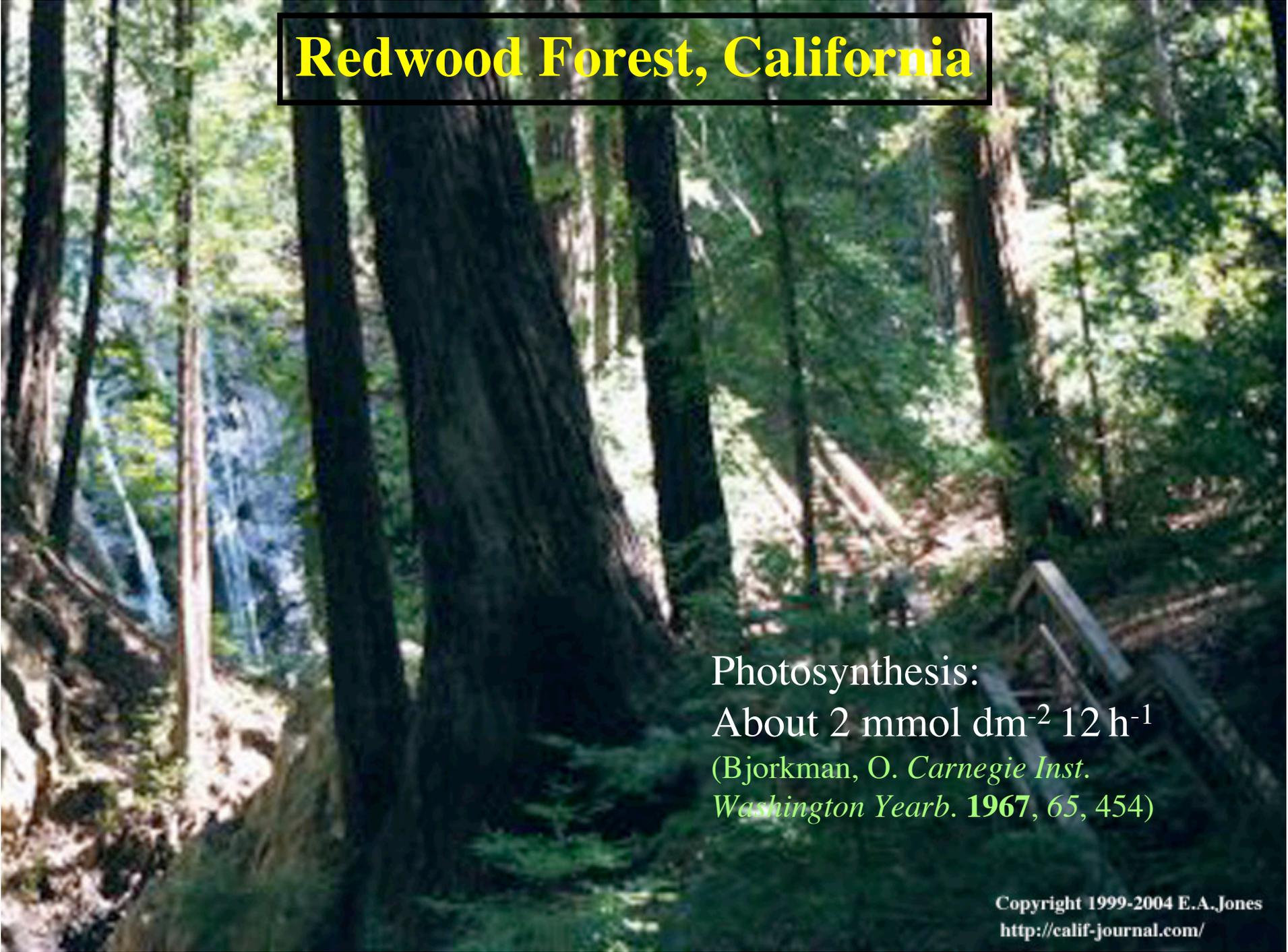
- 1) ISC'
- 2) bond



>97% Yield

Mortko, C. J.; Garcia-Garibay, M. A  
*J. Am. Chem. Soc.* **2005**, *127*, 7994

*How efficient is CO<sub>2</sub> fixation in nature?*



# Redwood Forest, California

Photosynthesis:  
About  $2 \text{ mmol dm}^{-2} 12 \text{ h}^{-1}$   
(Bjorkman, *O. Carnegie Inst.*  
*Washington Yearb.* **1967**, 65, 454)

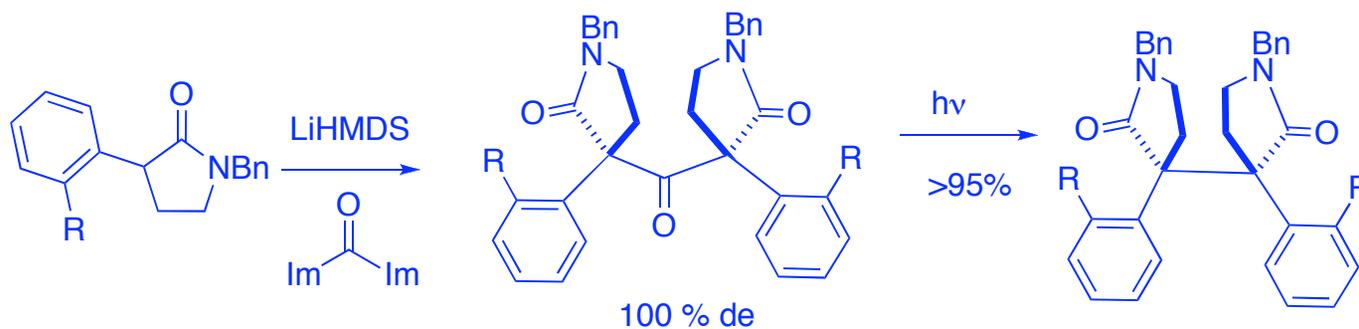
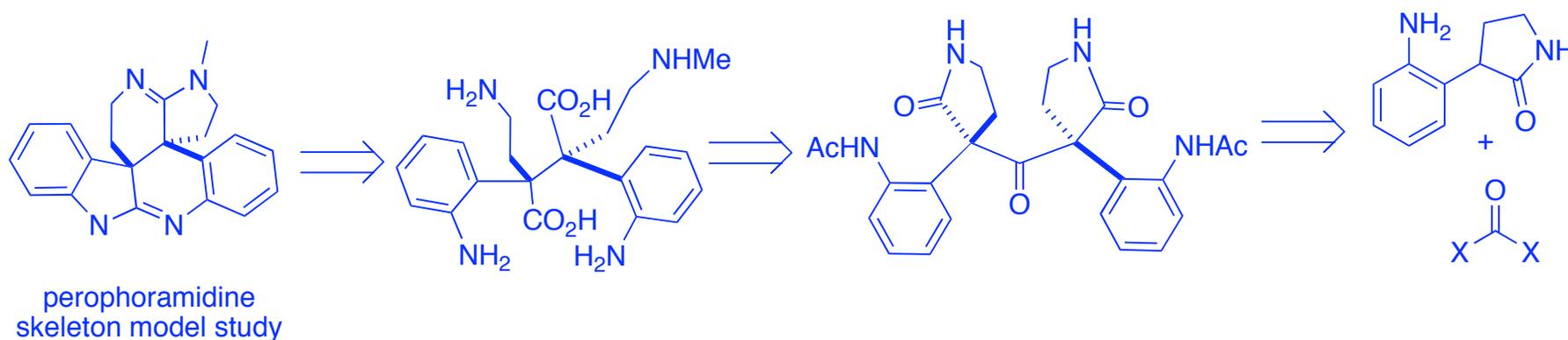
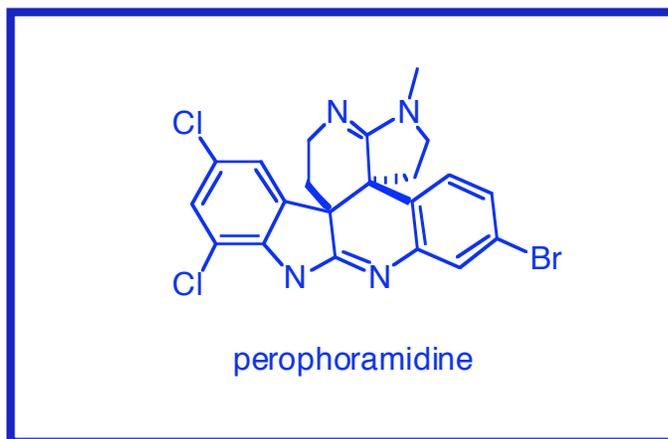


01/Feb/2005

2 x 100mg (0.75 mmol), 5h

=> ca. 0.9 mmol dm<sup>2</sup> 12 h<sup>-1</sup>)

# In Progress



# *Conclusions I*

## *Advantages:*

- a) A Challenging transformation made easy
- b) High selectivity
- c) High specificity
- d) No solvent

## *Disadvantages:*

- a) Must be and stay crystalline
- b) Highly substrate specific

*i.e., Very Much Like Most Enzymatic Processes!*  
*Potential Applications? Specialty Chemicals*

## Conclusions II

—*Solid state photochemical synthesis can be highly selective (there are many reactions amenable to development in the crystalline state)*

—*Nanocrystals make solid state reactions highly efficient*

—*Water suspensions make scale up easy to implement. Flow reactors are possible*

—*Complex structures, high yields, no workup (filtration), no purification*

### *Solar and Arc Lamp reactors*



**Precipitate  
and  
Filter**